

# Polyatomic Ions Chart

Common formulas, charges, and examples for introductory chemistry.

Ion	Formula	Charge	Examples
ammonium	NH <sub>4</sub>	+1	NH <sub>4</sub> Cl, (NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>
hydroxide	OH	-1	NaOH, Ca(OH) <sub>2</sub>
nitrate	NO <sub>3</sub>	-1	NaNO <sub>3</sub> , Ca(NO <sub>3</sub> ) <sub>2</sub>
nitrite	NO <sub>2</sub>	-1	NaNO <sub>2</sub>
sulfate	SO <sub>4</sub>	-2	Na <sub>2</sub> SO <sub>4</sub> , Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>
sulfite	SO <sub>3</sub>	-2	Na <sub>2</sub> SO <sub>3</sub>
carbonate	CO <sub>3</sub>	-2	CaCO <sub>3</sub> , Na <sub>2</sub> CO <sub>3</sub>
bicarbonate	HCO <sub>3</sub>	-1	NaHCO <sub>3</sub>
phosphate	PO <sub>4</sub>	-3	Na <sub>3</sub> PO <sub>4</sub> , Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>
acetate	C <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	-1	NaC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>
chlorate	ClO <sub>3</sub>	-1	KClO <sub>3</sub>
perchlorate	ClO <sub>4</sub>	-1	NaClO <sub>4</sub>
cyanide	CN	-1	NaCN
permanganate	MnO <sub>4</sub>	-1	KMnO <sub>4</sub>
chromate	CrO <sub>4</sub>	-2	K <sub>2</sub> CrO <sub>4</sub>
dichromate	Cr <sub>2</sub> O <sub>7</sub>	-2	K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub>